

Practice Test - Formulas and Naming Compounds

1. Write formulas for each of the following compounds

- | | |
|--------------------------------|---|
| a) Potassium oxide | <u>K₂O</u> |
| b) Strontium Bromide | <u>SrBr₂</u> |
| c) Sulfur tetroxide | <u>SO₄</u> |
| d) Lead IV carbonate | <u>Pb(CO₃)₂</u> |
| e) Phosphorus dichloride | <u>PCl₂</u> |
| f) Ammonium sulfate | <u>(NH₄)₂SO₄</u> |
| g) Iron II chloride trihydrate | <u>FeCl₂ · 3H₂O</u> |
| h) Sulfuric acid | <u>H₂SO₄</u> |
| i) Magnesium hydroxide | <u>Mg(OH)₂</u> |
| j) Nitric acid | <u>HNO₃</u> |

2. Name the following compounds

- | | |
|--|---|
| a) Fe ₂ O ₃ | <u>Iron (III) oxide</u> |
| b) Cs ₂ S | <u>Cesium sulfide</u> |
| c) N ₂ O ₅ | <u>Dinitrogen pentoxide</u> |
| d) HBr | <u>Hydrobromic acid</u> |
| e) CuCl ₂ · 5H ₂ O | <u>Copper (II) chloride pentahydrate</u> |
| f) CH ₃ COOH | <u>Acetic acid</u> |
| g) ZnSO ₄ | <u>Zinc (II) sulfate <u>or</u> Zinc sulfate</u> |

3. Give an example of a binary compound. K₂O

4. Give an example of an oxy-acid H₃PO₄

5. What is an anhydrous compound? Hydrates which have been dehydrated.