

Key

Polyatomic Ions

- polyatomic ions most commonly end in ate, but can also end in ite or ide
- do not ever split up the atoms in polyatomic ions, they stay together as a package (think of them as a single ion)

Write the formula for the following polyatomic ions (don't forget the charge)

- a) carbonate CO_3^{2-} f) hypochlorite ClO^-
b) nitrite NO_2^- g) bisulphide HS^-
c) oxalate $\text{C}_2\text{O}_4^{2-}$ h) hydroxide OH^-
d) acetate CH_3COO^- i) permanganate MnO_4^-
e) cyanide CN^- j) thiocyanate SCN^-

Write the name for the following polyatomic ions (underline the ending)

- a) ClO_3^- Chlorate f) $\text{Cr}_2\text{O}_7^{2-}$ Dichromate
b) HCO_3^- Bicarbonate g) ClO_4^- Perchlorate
c) SO_4^{2-} Sulfate h) HSO_3^- Bisulphite
d) NO_3^- Nitrate i) HC_2O_4^- Binoxalate
e) PO_4^{3-} Phosphate j) H_2PO_4^- Dihydrogen phosphate