

## Polyatomic Ions

- polyatomic ions most commonly end in ate, but can also end in ite or ide
- do not ever split up the atoms in polyatomic ions, they stay together as a package (think of them as a single ion)

Write the formula for the following polyatomic ions (don't forget the charge)

- a) carbonate \_\_\_\_\_ f) hypochlorite \_\_\_\_\_  
b) nitrite \_\_\_\_\_ g) bisulphide \_\_\_\_\_  
c) oxalate \_\_\_\_\_ h) hydroxide \_\_\_\_\_  
d) acetate \_\_\_\_\_ I) permanganate \_\_\_\_\_  
e) cyanide \_\_\_\_\_ j) thiocyanate \_\_\_\_\_

Write the name for the following polyatomic ions (underline the ending)

- a)  $ClO_3^-$  \_\_\_\_\_ f)  $Cr_2O_7^{2-}$  \_\_\_\_\_  
b)  $HCO_3^-$  \_\_\_\_\_ g)  $ClO_4^-$  \_\_\_\_\_  
c)  $SO_4^{2-}$  \_\_\_\_\_ h)  $HSO_3^-$  \_\_\_\_\_  
d)  $NO_3^-$  \_\_\_\_\_ I)  $HC_2O_4^-$  \_\_\_\_\_  
e)  $PO_4^{3-}$  \_\_\_\_\_ j)  $H_2PO_4^-$  \_\_\_\_\_